Honors Chemistry Hour\_\_\_\_\_ Name\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_
Dr. Wexler
Completing and Balancing Double Replacement Reactions

Date\_\_\_\_\_

Instructions: Use the Periodic Table of Ions as well as your understanding of ionic compounds to complete and balance each of the following chemical equations involving double replacement reactions:

1) \_\_\_ NaOH + \_\_\_ CaBr2 🡪

2) \_\_\_ Pb(NO3)2 + \_\_\_ HCl 🡪

3) \_\_\_ Na2CO3 + \_\_\_ KF 🡪

4) \_\_\_ AgNO3 + \_\_\_ CuSO4 🡪

5) \_\_\_ AgF + \_\_\_ NiCl2 🡪

6) Devise a method using double replacement for making silver chromate. You may invent any two ionic compounds as reactants providing that they produce two products, one of which is silver chromate, which is insoluble. Write the balanced equation below: