Honors Chemistry Hour\_\_\_\_\_ Name\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_  
Dr. Wexler  
Stoichiometry Worksheet 3 (HS-PS1-7)  
Date\_\_\_\_\_

Given the following reactants: copper(II) sulfate + sodium hydroxide

1. Write and balance the equation for the resulting double replacement reaction. The precipitate that forms is copper(II) hydroxide.

2. How many grams of copper(II) hydroxide will be produced from 200g of NaOH in excess CuSO4?  
  
Neatly organize your calculations into three sections using the MaMoMoMa Plan. Be sure to write the unit after each number in this process . Show all calculations!

A. Mass to Mole  
 - molar mass of NaOH

-convert mass to moles

B. Mole to Mole  
 -use stoichiometry to determine moles of copper(II) hydroxide

C. Mole to Mass  
 - molar mass of copper(II) hydroxide

-convert moles to mass