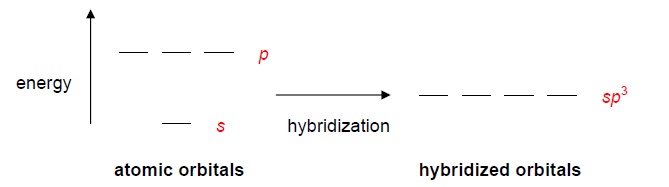
Honors Chemistry Hour\_\_\_\_\_ Name\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_  
Dr. Wexler  
Hybrid Atomic Orbitals  
Date\_\_\_\_\_

When atoms bond to form molecules, they use molecular orbitals. These are formed through the hybridization of the atomic orbitals *s*, *p*, and *d.*

The hybridized molecular orbitals have different shapes and energy levels than the atomic orbitals. The number of **molecular** orbitals created by hybridization depends on the number of **atomic** orbitals that are mixed to form them.

**I. sp3 hybridized orbitals:  
A. Carbon (when forming carbon tetrachloride, CCl4)**

In forming ***sp*3** hybridized orbitals, **four** atomic orbitals are mixed, one *s* and three *p*. The energy diagram for this process is shown below. The hybridized orbitals are higher in energy than the *s* orbital, but lower in energy than the *p* orbitals. Below is an incomplete electron orbital energy diagram for carbon:



Carbon has 4 valence electrons. Add these electrons to the atomic orbitals on the left and also distribute them in sp3 hybridized molecular orbitals on the right. With this hybridization, C will form **four equivalent** σ **bonds**. This hybridization gives the molecule a **tetrahedral geometry.**

Draw the Lewis diagram for carbon tetrachloride:

How many σ bonds are formed?

How many lone pairs of electrons does carbon have?

Draw a tetrahedron:

**B. Oxygen (when forming water, H2O)**  
Draw the electron orbital energy diagram for ***sp*3** hybridized **oxygen**, which has 6 valence electrons (refer to the carbon energy diagram above):

Draw the Lewis diagram for water

How many σ bonds are formed?

How many lone pairs of electrons does oxygen have?

How do the lone electron pairs affect the geometry of the water molecule?

**C. Nitrogen (when forming ammonia, NH3)** Draw the electron orbital energy diagram for ***sp*3** hybridized **nitrogen**, which has 5 valence electrons. Remember that, as with atomic orbitals, every hybridized molecular orbital must first be filled with one electron before electrons can be paired (Hund’s Rule).

Draw the Lewis diagram for ammonia:

How many σ bonds are formed?

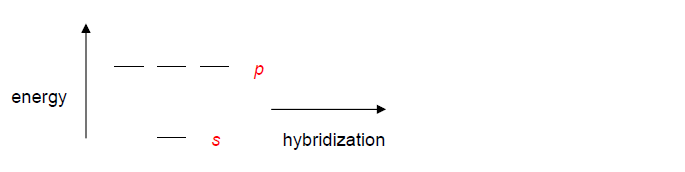
How many lone pairs of electrons does oxygen have?

How do the lone electron pairs affect the geometry of the molecule?

**II. sp2 hybridized orbitals:  
Al (when forming AlCl3)**

In some Lewis structures, only **three** equivalent bonds are formed. To create three equivalent hybridized sp2 orbitals, mix **three** atomic orbitals (one s and two p).

Complete the energy diagram below showing orbital hybridization in **aluminum**.

  
  
Draw the Lewis diagram for aluminum trichloride:

How many σ bonds are formed?

How many lone pairs of electrons does aluminum have?

This hybridization gives **\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_** geometry.

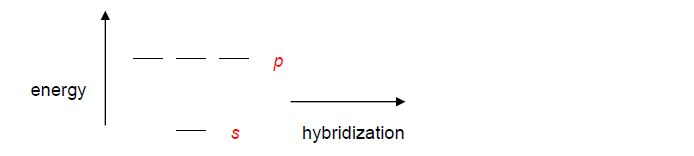
What is unusual about the number of electrons around aluminum in aluminum trichloride? What rule does this violate?

Calculate the formal charge on aluminum in this molecule (show your calculations). How does this help explain why the aluminum is stable?

**III. sp hybridized orbitals  
Be (when forming BeCl2)**

In some Lewis structures, only **two** equivalent bonds are formed. To create two equivalent hybridized sp orbitals, mix **two** atomic orbitals (one s and one p).

Complete the energy diagram below showing orbital hybridization in **beryllium.**



Draw the Lewis diagram for beryllium dichloride:

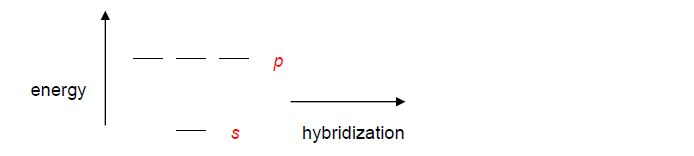
How many σ bonds are formed?

How many lone pairs of electrons does beryllium have?

This hybridization gives **\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_** geometry.

**IV. sp hybridized orbitals forming double bonds:**  
In **linear** molecules, like **CO2**, **two** equivalent hybridized sp orbitals are formed. To create two equivalent hybridized sp orbitals, mix **two** atomic orbitals (one s and one p). Note that two electron-containing p orbitals remain unhybridized.

Complete the energy diagram below showing orbital hybridization in **carbon**.



Draw the Lewis diagram for carbon dioxide:

How many σ bonds are formed?

How many π bonds are formed?

**Advanced:** Carefully diagram and label the orbitals and bonds of CO2