Honors Chemistry Hour\_\_\_\_\_ Name\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_
Dr. Wexler
Basic Atomic Structure Practice Quiz 1
Date assigned:

Answer the following questions:

1. What is the charge and mass of a single electron?
2. What are the two subatomic particles found in the nucleus of an atom?
3. Is the nucleus positively or negatively charged?
4. What subatomic particle determines the charge of the nucleus?
5. If a neutral atom has 7 protons, how many electrons does it have?
6. The sodium cation Na+1 has how many electrons?
7. The oxygen anion O-2 has how many electrons?
8. List the three subatomic particles:
A.
B.
C.
9. Which two of the subatomic particles account for almost all of an atom’s mass?
A.
B.
10. An isotope of Lead has 82 protons and 126 neutrons. What is the mass number of this isotope?
11. Given the isotope Uranium -238:
A. How many protons does it have?
B. How many electrons does it have?
C. How many neutrons does it have?
12. Chlorine, which has an atomic mass of 35.453 amu, has two naturally occurring isotopes, Cl-35 and Cl-37. Which of these two isotopes is more common? How do you know?